Roger Adams and H. W. Stewart. Restricted Rotation in Arylamines. II. Preparation and Resolution of N-Succinyl-N-ethyl-3-bromomesidine and 5-Alkoxy-4-N-succinyl-4-amino-1,3-dimethylbenzenes.

Page 2860. Col. 1, line 1 of last paragraph, for "(1 mole)" read "(0.1 mole)," and same below for diethyl sulfate and sodium nitrate.

T. Q. Chou and T. T. Chu. The Preparation and Properties of Peimine and Peiminine.

Page 2936. Col. 1, line 5, for "wereas signed" read "were assigned."

George Wash, Billie Shive and H. L. Lochte. Normal and Abnormal Alkylation of 2-Methylcyclopentyl Methyl Ketone.

Page 2975. In the title, for "Methyl" read "Phenyl."—H. L. LOCHTE.

W. D. Larson and W. J. Tomsicek. The Activity Coefficients of the Undissociated Part of Weak Acids. II. Oxalic Acid.

Page 3330. Equation (4) should read

$$E^0 = E + 0.02957 \log_{10} K_1 K_2 C \left(1 - \frac{-K_1 + \sqrt{K_1^2 + 4K_1 C}}{2C}\right)$$

—W. D. LARSON.

1942, VOL. 64

W. R. Forsythe and W. F. Giauque. The Entropies of Nitric Acid and its Mono- and Tri-hydrates. Their Heat Capacities from 15 to 300°K. The Heats of Dilution at 298.1°K. The Internal Rotation and Free Energy of Nitric Acid Gas. The Partial Pressures over its Aqueous Solutions.

Pages 60 and 61. The authors call attention to the following arithmetical errors:

$$^1/_2N_2 + {}^3/_2O_2 + {}^1/_2H_2 = HNO_3 \quad (1)$$

$\Delta F^0_{298.1}$ should be −19030 cal. instead of −11,539 cal.

$$^1/_2N_2 + {}^3/_2O_2 + {}^1/_2H_2 = HNO_3 \quad (g)$$

$\Delta F^0_{298.1}$ should be −17948 cal. instead of −10.457

$$^1/_2N_2 + {}^3/_2O_2 + {}^1/_2H_2 = HNO_3 \quad (a = 1)$$

$\Delta F_{298.1}$ should be −26345 cal. instead of −18854 cal.

The above errors are repeated in the summary.
Also in the summary

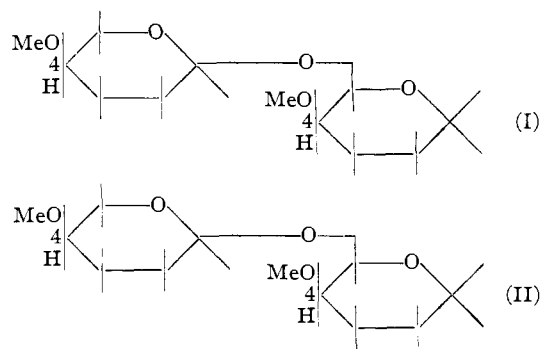
$$^1/_2N_2 + {}^3/_2O_2 + {}^1/_2H_2 = HNO_3 \quad (g)$$

$\Delta H_{298.1}$ should be −31994 cal. instead of −3199 cal.

—W. R. FORSYTHE and W. F. GIAUQUE.

E. V. White. The Constitution of Arabo-galactan. II. The Isolation of Heptamethyl- and Octamethyl-6-galactosidogalactose through Partial Hydrolysis of Methylated Arabo-galactan.

Page 303. The author writes "Formulas 1 and 2 representing octamethyl-6-*d*-galactosidogalactose and heptamethyl-6-*d*-galactosidogalactose are incorrect in the relative positions of the methoxyl and hydrogen groups at position 4 in all the monosaccaride units, which should be reversed as indicated



Page 304. In col. 1 of the Experimental, line 30, in the formula, for "Hg" read "H_8."—E. V. WHITE.

Emmett R. Barnum and Cliff S. Hamilton. 5-Amino- and 1-Aminobenzo(f)-quinolines and Derivatives.

Page 542. Col. 2, lines 11 and 20, for "*p*-dimethylaminobenzaldehyde" read "*p*-diethylaminobenzaldehyde."—C. S. HAMILTON.

H. J. Lucas and Clark W. Gould, Jr. Brucine as a Reagent for Partially Resolving Bromoalkanes; the Configurations of Some Diastereomeric Dibromoalkanes.

Page 602. In footnote (7) for "(+1.0016)" read "(0.0016)," for "compound" read "compared" and in the citation, for "1471" read "147."

N. Howell Furman and Clark E. Bricker. A Polarographic Study of *o*-Phthalic Acid and Phthalates.

Page 665. In the legend for Fig. 6, the last four lines should read "Curve a at *p*H 4.02; curve b at *p*H 3.94; curve c at *p*H 3.75; curve d at *p*H 3.57; curve e at *p*H 3.48; curve f at *p*H 3.25; curve g at *p*H 3.06; curve h at *p*H 2.60; polarogram was made with capillary no. 3."—N. H. FURMAN.

S. C. Schumann, J. G. Aston and Malcolm Sagenkahn. The Heat Capacity and Entropy, Heats of Fusion and Vaporization and the Vapor Pressures of Isopentane.

Page 1041. The authors write "In Table IV the molal heat of vaporization of isopentane at 298.16°K. calculated from Eq. (1) and the modified Berthelot equation at 298.16°K. with T_c = 461°K. and P_c = 32.9 atm. is 5965 calories (Berthelot correction = 239 calories), instead of the value given."

Page 1043. "In Table VII the free energy of neopentane was accidentally computed using the free rota-